

DEC 19

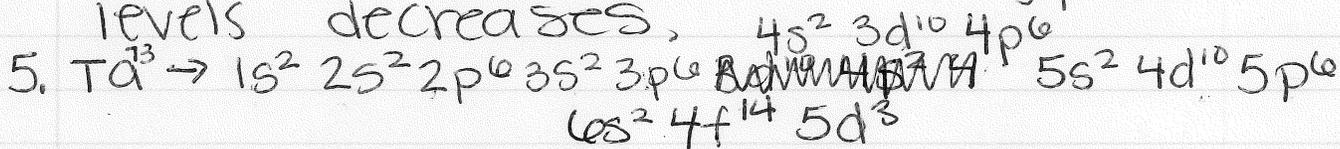
Grau 3 Boulton Chapter 5

1. RUTHERFORD'S model didn't explain electrons BEHAVIOR OR chemical properties.  
↳ moving to different energy level rings
2. THOMPSON model: atoms "pudding" where electrons are evenly spread out around the nucleus

Bohr: nucleus in the middle rings (electrons)  
- shows where electrons might be

3. Quanta

4. increases however the energy to move levels decreases.



6. As one increases the other decreases

7. Wavelength & length of one full wave

Frequency: how many full waves pass at a given time

8.  $c = \lambda \nu$

$c = 2.998 \cdot 10^8 \text{ m/s}$

$\lambda = 7 \text{ m}$

$\nu = ? 42828571.43 \text{ s}$

~~$e = h\nu$~~

$e = ? = 2.837 \times 10^{-26}$

$h = 6.626 \cdot 10^{-34}$

$\nu = 42828571.43$

9. Radio MICROWAVE Infrared ~~Ultraviolet~~

Ultraviolet XRay

Gamma

10. Electrons go from higher rings to lower rings.